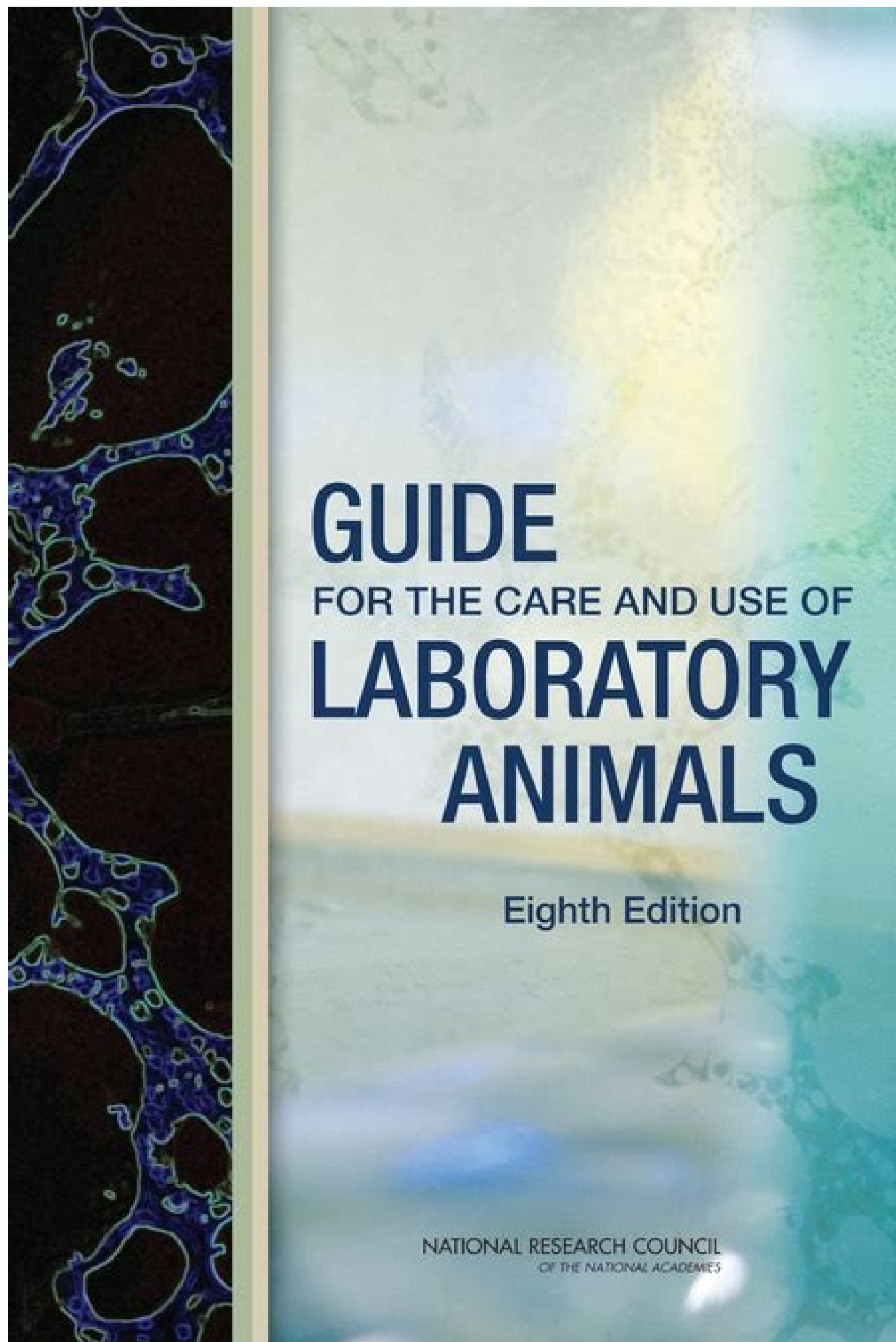


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**\*Abstract:**

Alkalinity is the ability of water to resist change in PH, which is due to presence of certain species such as hydroxide ions  $\text{OH}^{-1}$ , bicarbonate ions  $\text{HCO}_3^{-1}$  and carbonate ions  $\text{CO}_3^{-2}$ .  
In this experiment we will calculate

$$\text{Total Alkalinity} = [\text{HCO}_3^{-1}] + 2[\text{CO}_3^{-2}] + [\text{OH}^{-1}] - [\text{H}^{+}]$$

$$\text{Carbonate ALK} = [\text{CO}_3^{-2}] + [\text{OH}^{-1}] - [\text{H}^{+}] - [\text{H}_2\text{CO}_3]$$

$$\text{Caustic ALK} = [\text{OH}^{-1}] - [\text{H}^{+}] - [\text{HCO}_3^{-1}] - 2[\text{H}_2\text{CO}_3]$$

The alkalinity of water can be determined by titrating the water sample with Sulphuric acid of known values of pH, volume and concentrations. Based on stoichiometry of the reaction and number of moles of Sulphuric acid needed to reach the end point, the concentration of alkalinity in water is calculated.

**\*Introduction:**

Alkalinity is primarily a way of measuring the acid neutralizing capacity of water. In other words, its ability to maintain a relatively constant pH. The possibility to maintain constant pH is due to the hydroxyl, carbonate and bicarbonate ions present in water. The ability of natural water to act as a buffer is controlled in part by the amount of calcium and carbonate ions in solution. Carbonate ion and calcium ion both come from calcium carbonate or limestone. So water that comes in contact with limestone will contain high levels of both  $\text{Ca}^{++}$  and  $\text{CO}_3^{-2}$  ions and have elevated hardness and alkalinity. In the other hand the **ENVIRONMENTAL SIGNIFICANCE occur as following**

Alkalinity is important for fish and aquatic life because it protects or buffers against rapid pH changes. Higher alkalinity levels in surface waters will buffer acid rain and other acid wastes and prevent pH changes that are harmful to aquatic life.

Large amount of alkalinity imparts bitter taste in water.

The principal objection of alkaline water is the reactions that can occur

